

Free reading Synthesis analysis of aspirin lab answers (Download Only)

study with quizlet and memorize flashcards containing terms like synthesis of aspirin acetic anhydride two step reaction and more section 1 purpose and summary conduct a chemical reaction to produce aspirin separate the aspirin from the reaction by products using vacuum filtration analyze the aspirin and estimate its purity acetylsalicylic acid commonly known as aspirin is the most widely used drug in the world today again for aspirin the melting point is 135 degree celsius and this data falls in the range of the obtained experimental data therefore the product contains pure aspirin tlc the below tlc analysis showed that the acetylsalicylic had a higher rf value than salicylic acid 0 vs 0 rf value respectively which proves that salicylic acid was what is the theoretical yield of aspirin in two significant figures why is the aspirin washed with cold water according to the data in the merck index if 1.0 g of aspirin is dissolved in 100 ml of water at 37 c how much aspirin will precipitate out of solution when it is cooled to 25 c the synthesis of aspirin is an organic chemistry experiment in many specifications for students of ages 16 18 years each of the four levels take approximately 30 minutes to complete and are designed to be used as pre lab activities in class or as homework these explanatory demonstration lesson 180 minutes day 1 synthesis of aspirin lab activity students should read the lab and answer the pre lab questions prior to arrival tutorial on how to complete the aspirin synthesis b virtual lab level 1 website link virtual.edu/rsc.org note this is not really an answer key do n synthesis of aspirin b lab purpose to synthesize aspirin b and determine its purity using the ferric chloride test to determine aspirin b percentage yield data figure 1 balanced aspirin b synthesis reaction mass of aspirin b calculation molar mass of salicylic acid 138 g/mol mass of salicylic acid 3 g 138 g/mol 0 moles synthesis of aspirin by jon torre purpose to determine which of four catalysts yields the fastest reaction rate in the acetylation of salicylic acid 1 to form acetylsalicylic acid 2 reactions procedure and results aspirin synthesis tap water was heated on a steam bath in a 250 ml beaker the objective of this experiment is to enable us to conduct the synthesis of aspirin reinforce the skills of recrystallisation and reinforce the technique of melting point determination it is carried out to form ester from an acid and an alcohol the main procedures are preparation of aspirin recrystallisation of aspirin and lastly aspirin synthesis and analysis revised 12 13 14 crush an aspirin tablet and place in a labeled pre weighed vial or test tube record the mass of the aspirin tablet do not lose this container or its contents you will use them later take a small amount of the aspirin tablet and dissolve as much as you can in 3 drops of ethyl acetate 1 measure out 2.0 g of salicylic acid sa and place the sa in to a 50 ml erlenmeyer flask 2 measure out 5.0 ml of acetic anhydride d 1.08 g/ml in a graduated cylinder and then pour it into the flask in such a way as to wash any crystals of sa on the walls down to the bottom 3 add 5 drops of 85 phosphoric acid to serve as a catalyst transfer the impure aspirin from the büchner funnel to a 150 ml beaker add 10 ml of 95 ethanol add about 25 ml of warm 50-70 c distilled water and stir to dissolve the impure aspirin the mixture can be heated gently until the solid completely dissolves but do not allow the mixture to boil in this lab we will be using a chemical test to determine the purity of the aspirin you synthesized this particular chemical test takes advantage of one of the differences between salicylic acid and aspirin namely the functional group called phenol a functional group is a group of atoms that have characteristic reactivity the purpose of this experiment is to synthesize aspirin the same way performed by manufacturers this will be accomplished by adding acetic anhydride into salicylic acid heating sulphuric acid and determining the reactions based on the colors of the solutions created at the end of the experiment

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section 1 purpose and summary conduct a chemical reaction to produce aspirin separate the aspirin from the reaction by products using vacuum filtration analyze the aspirin and estimate its purity acetylsalicylic acid commonly known as aspirin is the most widely used drug in the world today

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the synthesis of aspirin is an organic chemistry experiment in many specifications for students of ages 16 18 years each of the four levels take approximately 30 minutes to complete and are designed to be used as pre lab activities in class or as homework these explanatory demonstration

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synthesis of aspirin b lab purpose to synthesize aspirin b and determine its purity using the ferric chloride test to determine aspirin b percentage yield data figure 1 balanced aspirin b synthesis reaction mass of aspirin b calculation molar mass of salicylic acid 138 g mol mass of salicylic acid 3 g 138 g mol 0 moles

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synthesis of aspirin by jon torre purpose to determine which of four catalysts yields the fastest reaction rate in the acetylation of salicylic acid 1 to form acetylsalicylic acid 2 reactions procedure and results aspirin synthesis tap water was heated on a steam bath in a 250 ml beaker

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the objective of this experiment is to enable us to conduct the synthesis of aspirin reinforce the skills of recrystallisation and reinforce the technique of melting point determination it is carried out to form ester from an acid and an alcohol the main procedures are preparation of aspirin recrystallisation of aspirin and lastly

aspirin synthesis analysis

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aspirin synthesis and analysis revised 12 13 14 crush an aspirin tablet and place in a labeled pre weighed vial or test tube record the mass of the aspirin tablet do not lose this container or its contents you will use them later take a small amount of the aspirin tablet and dissolve as much as you can in 3 drops of ethyl acetate

experiment synthesis of aspirin bellevue college

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1 measure out 2.0 g of salicylic acid sa and place the sa in to a 50 ml erlenmeyer flask 2 measure out 5.0 ml of acetic anhydride d 1.08 g/ml in a graduated cylinder and then pour it into the flask in such a way as to wash any crystals of sa on the walls down to the bottom 3 add 5 drops of 85 phosphoric acid to serve as a catalyst

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transfer the impure aspirin from the büchner funnel to a 150 ml beaker add 10 ml of 95 ethanol add about 25 ml of warm 50 70 c distilled water and stir to dissolve the impure aspirin the mixture can be heated gently until the solid completely dissolves but do not allow the mixture to boil

83 experiment chem21labs

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in this lab we will be using a chemical test to determine the purity of the aspirin you synthesized this particular chemical test takes advantage of one of the differences between salicylic acid and aspirin namely the functional group called phenol a functional group is a group of atoms that have characteristic reactivity

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